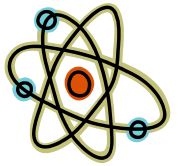
Atom: The smallest particle into which n element can be divided and still have the properties of that element.

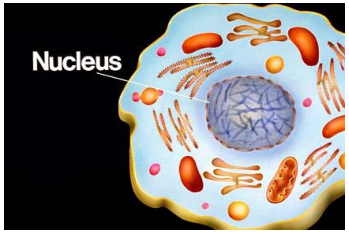
Proton: A positively- charged particle that is located in the nucleus of the atom (+) They can be identified by looking at the atomic number.

Neutron: A particle with no charge that is located in the nucleus of the atom. (0) Mass # - Atomic # = Neutrons.

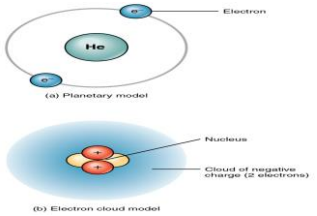
Electron: A negatively- charged particle that is located outside the nucleus of the atom. (-) There are the same number of protons and electrons in an atom and can be identified by the Atomic #.



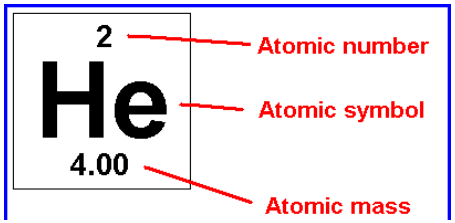
Nucleus: The center of an atom, containing the protons and neutrons. It has on overall charge of (+).



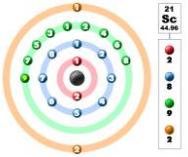
Electron Cloud: The region of the atom where electrons may be found. It has an overall charge of of (-).



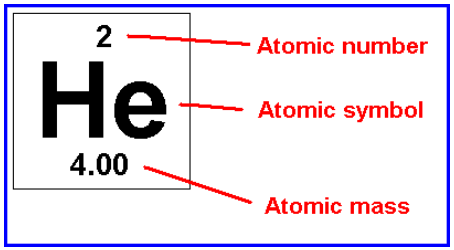
Atomic Number: The number of Protons in the Nucleus of an atom. Atomic number identifies the element.
Atomic # = Protons = Electrons



Energy level/ Orbital: A specific Area in an Atom at a definite distance from the nucleus where electrons are arranged.
1st Orbital holds 2 electrons, 2nd orbital holds 8, 3rd orbital holds 18 (but is happy with 8)

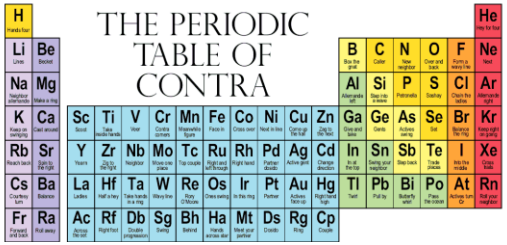


Mass Number: The sum of the protons and Neutrons.
Mass Number – Atomic Number (protons) = Neutrons.

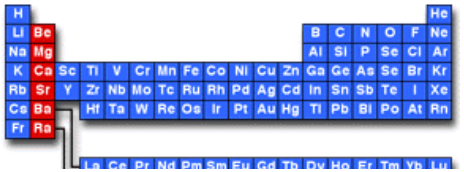


Alkali Metals- Elements that form alkali solutions when they combine with water. These elements have one valence electron, and therefore react easily with other elements and are found in nature only in compounds. These elements are in group 1 on the periodic Table.

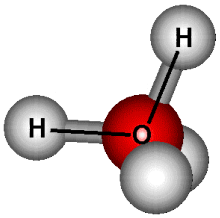
Element- A pure substance that cannot be broken down into simpler substance.



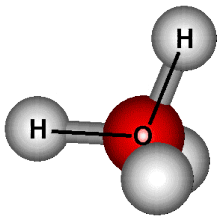
Alkaline Earth Metals: The elements in group 2 on the Periodic Table. These elements are very reactive but are less reactive that alkali metals; their atoms have two valence electrons.



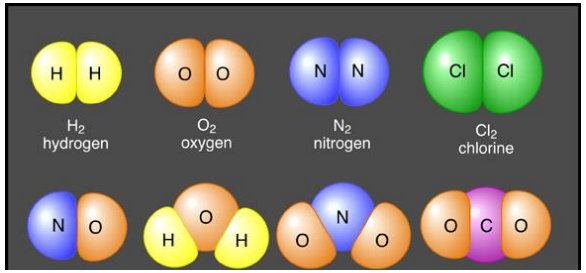
Molecule: A substance made of two or more elements that are chemically combines.



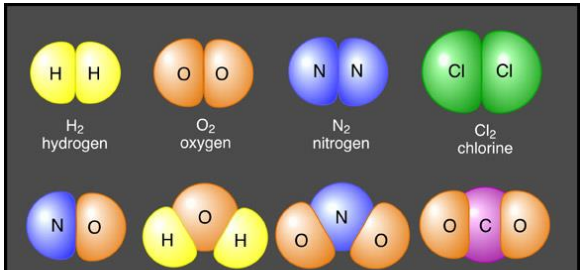
Molecule: A substance made of two or more elements that are chemically combines.



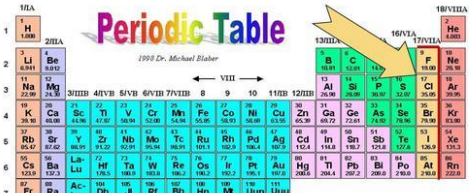
Compound: Two or more bonded atoms.



Compound: Two or more bonded atoms.



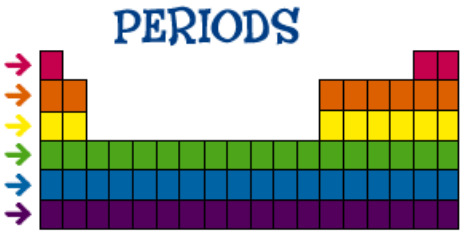
Halogens: Element that are in group 17 on the Periodic Table. These elements are very reactive; their atoms have seven valence electrons.



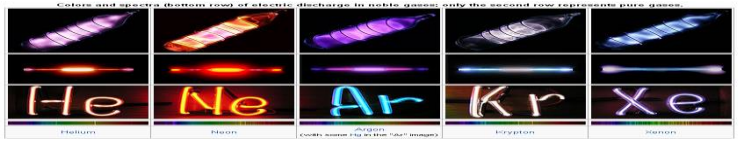
•Bending, Crushing, Mixing, Cutting, etc

•Change CANNOT be undone

Periods: A horizontal row of elements on the periodic table. It indicates the number if orbitals.

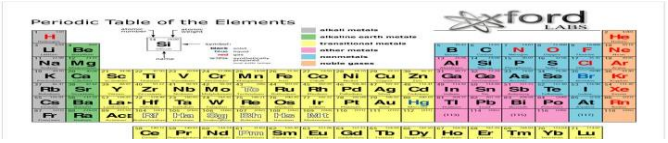


Nobel Gasses: Elements that are in group 18 of the Periodic Table. These elements are non-reactive nonmetals; their atoms have eight valence electrons (exceptions: Helium,- contains 2 valence electrons)



Physical property: A property of a substance that can be observed without changing the identity of the matter.

Group/ Families: A vertical column of elements on the periodic table. It indicates the number of valence electrons. Elements in the same column react similarly.



Chemical Property: A property of a substance that can be observed without changing the identity of the matter.